Chemistry SAT Practice Test 27

Q1. May be used in combination with a calorimeter to compare the specific heats of two substances
A. Thermometer
B. Conductivity tester
C. Salt bridge
D. Buret
E. Graduated cylinder

Q2. Is used to measure the volume of a solid by water displacement
A. Thermometer
B. Conductivity tester
C. Salt bridge
D. Buret
E. Graduated cylinder

Q3. Useful for adding small quantities of acid into a base
A. Thermometer
B. Conductivity tester
C. Salt bridge
D. Buret
E. Graduated cylinder

Q4. Completes the circuit of an electrochemical cell
A. Thermometer
B. Conductivity tester
C. Salt bridge
D. Buret
E. Graduated cylinder

Q5. Always amphoteric in nature
A. Nucleic acids
B. Proteins
C. Carbohydrates
D. Lipids
E. Electrolytes

Q6. Found as both straight-chained and branched polymers
A. Nucleic acids
B. Proteins
C. Carbohydrates
D. Lipids
E. Electrolytes

Q7. Deoxyribose in DNA nucleotides belongs to this family of biologically important molecules
A. Nucleic acids
B. Proteins
C. Carbohydrates
D. Lipids
E. Electrolytes

Q8. Always ionic in nature
A. Nucleic acids
B. Proteins
C. Carbohydrates
D. Lipids
E. Electrolytes

Q9. Tend not to be water soluble, and aggregate into droplets or molecular bilayers
A. Nucleic acids
B. Proteins
C. Carbohydrates
D. Lipids
E. Electrolytes

Q10. Represents the decomposition of a compound into its constituent elements
A. $Ag^+ + Br^- \rightarrow AgBr$
B. ${}^{14}_{6}C \rightarrow {}^{14}_{7}N + {}^{0}_{-1}e$
C. ${}^{234}_{92}U \rightarrow {}^{230}_{90}Th + {}^{4}_{2}He$
D. $+ {}^{30}_{15}P \rightarrow {}^{30}_{14}Si + {}^{0}_{1}e$
E. $2HgO \rightarrow 2Hg + O_2$

Q11. Represents alpha decay
A. $Ag^+ + Br^- \rightarrow AgBr$
B. ${}^{14}_{6}C \rightarrow {}^{14}_{7}N + {}^{0}_{-1}e$
C. ${}^{234}_{92}U \rightarrow {}^{230}_{90}Th + {}^{4}_{2}He$
D. $+ {}^{30}_{15}P \rightarrow {}^{30}_{14}Si + {}^{0}_{1}e$
E. $2HgO \rightarrow 2Hg + O_2$

Q12. Causes the neutron-to-proton ratio in a nucleus to be lowered

A. $Ag^{+} + Br^{-} \rightarrow AgBr$

B. ${}^{14}_{6}C \rightarrow {}^{14}_{7}N + {}^{0}_{-1}e$

C. ${}^{234}_{92}U \rightarrow {}^{230}_{90}Th + {}^{4}_{2}He$

D. $+ {}^{30}_{15}P \rightarrow {}^{30}_{14}Si + {}^{0}_{1}e$

E. $2HgO \rightarrow 2Hg + O_2$